

Counting atoms

Objectives:

1. To learn to count the amount of atoms in a compound
2. To learn that compounds are made of atoms from different elements

Procedures:

1. Count the amount of atoms of each elements in the compounds below
2. Write the amount of atoms of each element on the space provided

Examples:

<p>Sulfuric acid H₂SO₄</p> <p>Hydrogen (H) = 2 Sulfur () = Oxygen () =</p> <p style="text-align: center;">Total =</p>	<p>Carbon (C) = 8 x 2 = 16</p> <p>2 Caffeine molecules 2C₈H₁₀N₄O₂</p> <p>Hydrogen () = ___ x ___ = ___ (N) = ___ x 2 = ___ () = 2 x ___ = 4</p> <p style="text-align: center;">Total =</p>
<p>Ammonium acetates (NH₄)₂CH₃COO</p> <p>Nitrogen (N) = 2 (H) = Carbon () = (O) =</p> <p style="text-align: center;">Total =</p>	<p>Ammonium acetate 4(NH₄)₂CH₃COO</p> <p>(N) = ___ x 4 = ___ (H) = 11 x 4 = 44 () = ___ x ___ = ___ () = ___ x ___ = ___</p> <p style="text-align: center;">Total =</p>

Name and chemical formula	Amount of atoms
Water H ₂ O	
Hydrogen peroxide 2H ₂ O ₂	
Ammonium sulfide 2(NH ₄) ₂ S	
Sugar C ₆ H ₁₂ O ₆	

Name and chemical formula	Amount of atoms
Terephthalic acid $C_6H_4(COOH)_2$	
Ammonium $3NH_4HCO_3$	
Ammonium hydroxide NH_4OH	
Sodium acetate $NaCH_3COO$	
Halothane $C_2HBrClF_3$	
Acetic acid $2CH_3CO(OH)_2$	
Acetone $2(CH_3C)_3OCH_3$	