## **Identifying Unknown Elements**

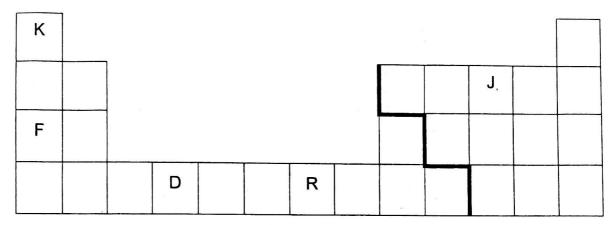
Use the periodic table to identify the elements described in the statement below. Write both the element name and symbol on the line provided. \_\_\_\_ 1. This element is in the same family as lead and has fewer protons than sodium.  $_{-}$  2. This element has an atomic number that is one greater than platinum, 3. This element has the most protons of any element in group 15. 4. This element has more than 50 but less than 75 protons, and it is in Group 17. 5. This Group 2 element has fewer protons than bromine but more protons than sulfur. 6. This element has the lowest atomic number of Group 16 elements. 7. This element has an atomic number that is double the atomic number of silicon. 8. This element has more valence electrons than oxygen, fewer valence electrons than neon, more protons than sodium but fewer protons than argon. 9. This element has an atomic number lower than that of aluminum and one less valence electron than the Group 16 elements. 10. This element is in Group 1 and has a higher atomic number than chlorine

but a lower atomic number than bromine.

Name	Date	Period

## **Interpreting the Periodic Table**

Examine the hypothetical periodic table shown below. Use this periodic table to answer the questions that follow.



- 1. Which pair of elements has the same number of valence electrons?
- 2. Which pair of elements is in the same period?
- 3. Which pair of elements is in the same group?
- 4. Which element has the smallest atomic number?
- 5. Which element(s) would be classified as a metal?
- 6. If the atomic number of Element D is 20, what is the atomic number of Element R?